

Ref #	Hits	Search Query	DBs	Default Operator	Plurals	Time Stamp
L1	0	("separationandguanidineandalkalimet al").PN.	USPAT; EPO; DERWENT	OR	OFF	2006/01/27 16:44
L2	0	separtion same guanidine	USPAT; EPO; DERWENT	AND	ON	2006/01/27 16:44
L3	29055	guanidine	USPAT; EPO; DERWENT	AND	ON	2006/01/27 16:44
L4	316	3 and "alkali metal halide"	USPAT; EPO; DERWENT	AND	ON	2006/01/27 16:45
L5	37	4 and separation and membrane	USPAT; EPO; DERWENT	AND	ON	2006/01/27 16:54
L6	1	5 and nanofiltration	USPAT; EPO; DERWENT	AND	ON	2006/01/27 16:46
L7	0	guanidine purification and carbonaceous adsorbent	USPAT; EPO; DERWENT	AND	ON	2006/01/27 16:55
L8	0	guanidine purification and activated carbon	USPAT; EPO; DERWENT	AND	ON	2006/01/27 16:55
L9	51	4 and activated carbon	USPAT; EPO; DERWENT	AND	ON	2006/01/27 16:59
L10	21	210/651 and 3	USPAT; EPO; DERWENT	AND	ON	2006/01/27 17:06
L11	54	separation of amine same guanidine	USPAT; EPO; DERWENT	AND	ON	2006/01/27 17:07
L12	13	11 and carbon and adsorption	USPAT; EPO; DERWENT	AND	ON	2006/01/27 17:11
L13	7	210/652 and guanidine	USPAT; EPO; DERWENT	AND	ON	2006/01/27 17:11

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Search Results - Record(s) 1 through 1 of 1 returned.

☐ 1. Document ID: US 5147553 A

L6: Entry 1 of 1

File: USPT

Sep 15, 1992

US-PAT-NO: 5147553

DOCUMENT-IDENTIFIER: US 5147553 A

TITLE: Selectively permeable barriers

DATE-ISSUED: September 15, 1992

INVENTOR-INFORMATION:

NAME	CITY	STATE	ZIP CODE	COUNTRY
Waite; Warren A.	Burlington	MA		

US-CL-CURRENT: [210/654](#); [210/490](#), [210/500.34](#), [427/245](#)

Full	Title	Citation	Front	Review	Classification	Date	Reference	References	Attachments	Claims	KWIC	Draw. Data
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L5 and nanofiltration	1

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Search Results - Record(s) 1 through 6 of 6 returned.

☐ 1. Document ID: US 5547817 A

L9: Entry 1 of 6

File: USPT

Aug 20, 1996

US-PAT-NO: 5547817

DOCUMENT-IDENTIFIER: US 5547817 A

TITLE: Photographic processing using a novel chelating complex

DATE-ISSUED: August 20, 1996

INVENTOR-INFORMATION:

NAME	CITY	STATE	ZIP CODE	COUNTRY
Okada; Hisashi	Kanagawa			JP
Inaba; Tadashi	Kanagawa			JP
Suzuki; Ryo	Kanagawa			JP
Nomura; Hideaki	Kanagawa			JP

US-CL-CURRENT: [430/393](#); [430/372](#), [430/428](#), [430/430](#), [430/461](#), [430/463](#), [430/488](#),
[430/490](#), [430/491](#)

Full	Title	Citation	Front	Review	Classification	Date	Reference	Amendments	Attachments	Claims	KIMC	Draw. Data
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☐ 2. Document ID: US 5447834 A

L9: Entry 2 of 6

File: USPT

Sep 5, 1995

US-PAT-NO: 5447834

DOCUMENT-IDENTIFIER: US 5447834 A

TITLE: Color diffusion transfer photographic material

DATE-ISSUED: September 5, 1995

INVENTOR-INFORMATION:

NAME	CITY	STATE	ZIP CODE	COUNTRY
Fujita; Munehisa	Kanagawa			JP
Asanuma; Hiroyuki	Kanagawa			JP

US-CL-CURRENT: [430/559](#); [430/217](#), [430/223](#), [430/569](#), [430/572](#), [430/574](#), [430/599](#),
[430/606](#)

Full	Title	Citation	Front	Review	Classification	Date	Reference	Sequences	Attachments	Claims	KWIC	Draw D
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☐ 3. Document ID: US 5250396 A

L9: Entry 3 of 6

File: USPT

Oct 5, 1993

US-PAT-NO: 5250396

DOCUMENT-IDENTIFIER: US 5250396 A

TITLE: Method for processing silver halide color photographic material

DATE-ISSUED: October 5, 1993

INVENTOR-INFORMATION:

NAME	CITY	STATE	ZIP CODE	COUNTRY
Ueda; Shinji	Kanagawa			JP
Ishikawa; Takatoshi	Kanagawa			JP
Fujimoto; Hiroshi	Kanagawa			JP

US-CL-CURRENT: 430/357; 430/398, 430/439, 430/484, 430/486, 430/490

Full	Title	Citation	Front	Review	Classification	Date	Reference	Sequences	Attachments	Claims	KWIC	Draw D
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☐ 4. Document ID: US 5147553 A

L9: Entry 4 of 6

File: USPT

Sep 15, 1992

US-PAT-NO: 5147553

DOCUMENT-IDENTIFIER: US 5147553 A

TITLE: Selectively permeable barriers

DATE-ISSUED: September 15, 1992

INVENTOR-INFORMATION:

NAME	CITY	STATE	ZIP CODE	COUNTRY
Waite; Warren A.	Burlington	MA		

US-CL-CURRENT: 210/654; 210/490, 210/500.34, 427/245

Full	Title	Citation	Front	Review	Classification	Date	Reference	Sequences	Attachments	Claims	KWIC	Draw D
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☐ 5. Document ID: US 4835092 A

L9: Entry 5 of 6

File: USPT

May 30, 1989

US-PAT-NO: 4835092

DOCUMENT-IDENTIFIER: US 4835092 A

TITLE: Method for processing a silver halide color photographic material using a color developer containing a chelating agent in the absence of benzyl alcohol

DATE-ISSUED: May 30, 1989

INVENTOR-INFORMATION:

NAME	CITY	STATE	ZIP CODE	COUNTRY
Ishikawa; Takatoshi	Kanagawa			JP
Koshimizu; Toshio	Kanagawa			JP
Fujita; Yoshihiro	Kanagawa			JP
Fujimoto; Hiroshi	Kanagawa			JP

US-CL-CURRENT: 430/380; 430/372, 430/434, 430/442, 430/464, 430/467, 430/468,
430/490, 430/491

Full	Title	Citation	Front	Review	Classification	Date	Reference	References	Attachments	Claims	KWIC	Draw D
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☐ 6. Document ID: US 4332733 A

L9: Entry 6 of 6

File: USPT

Jun 1, 1982

US-PAT-NO: 4332733

DOCUMENT-IDENTIFIER: US 4332733 A

TITLE: Process for liquefying acid anhydride

DATE-ISSUED: June 1, 1982

INVENTOR-INFORMATION:

NAME	CITY	STATE	ZIP CODE	COUNTRY
Samejima; Hiroshi	Ibaraki			JP
Miura; Mareki	Ibaraki			JP
Onuma; Yoshinobu	Ibaraki			JP

US-CL-CURRENT: 549/240; 549/255

Full	Title	Citation	Front	Review	Classification	Date	Reference	References	Attachments	Claims	KWIC	Draw D
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L8 and activated carbon

6

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L22: Entry 26 of 30

File: USPT

Jul 8, 1980

DOCUMENT-IDENTIFIER: US 4211880 A

TITLE: Resin-metal compound complex

Brief Summary Text (8):

Moore et al., in the Journal of the American Chemical Society, Vol. 72 (1950), pages 5792-3, disclose the adsorption of iron from a relatively concentrated hydrochloric acid solution on Dowex-1 by anion exchange.

Detailed Description Text (9):

The resin matrix of the weak base anion exchange resin contains chemically bonded thereto a basic, non-ionic functional group. The functional groups include primary, secondary, or tertiary amine groups. These may be aliphatic, aromatic, heterocyclic or cycloalkane amine groups. They may also be diamine, triamine, or alkanolamine groups. The amines may include alpha, alpha'-dipyridyl, guanidine, and dicyandiamidine groups. Other nitrogen-containing basic, non-ionic functional groups include nitrile, cyanate, isocyanate, thiocyanate, isothiocyanate, and isocyanide groups. Pyridine groups may also be employed.

Detailed Description Text (49):

Although low temperature liquid phase reactions are preferred, it will be appreciated that many low temperature reactions involve gaseous reactants and may be carried out in the gas phase, and the complexes can be applicable in these reactions. In some reactions, both liquid and gaseous reactants take part and are suitably catalyzed by the complexes. In all reactions, ease of catalyst separation by conventional operations of filtration, decantation, or centrifugation is a characteristic, whether the products and/or reactants are liquid or gaseous. The reactions may be carried out in conventional fixed bed flow reactors, or in continuously stirred flow reactors, or in batch reactors. Pressure may range to 300 atmospheres or more, and reaction times from less than one minute to several hours.

Detailed Description Text (69):

In place of KCl, other alkali metal halides, alkaline earth metal halides, especially the chlorides, and ammonium chloride, may be used.

Detailed Description Text (80):

The reaction mixture was heated to the reflux temperature to melt the sodium which was dispersed by the high speed stirrer. The resultant sodium dispersion was then cooled to room temperature and to it were added 10 ml. of a solution composed of 37.2 g. of chlorodiphenylphosphine dissolved in 50 ml. of para-dioxan. The resultant mixture was heated to the reflux temperature and 0.5 g. of naphthalene was added. The remainder of the chlorodiphenylphosphine solution was added to the reaction mixture dropwise over a period of three hours and the final reaction mixture was maintained at the reaction temperature for an additional 16 hours. The solution was cooled at 60.degree. C. and a 50 ml. aliquot of the reaction solution was withdrawn. To the remaining reaction mixture was added 50 g. of a solid resin, the resin being a copolymer of styrene and divinylbenzene which contained chloromethyl functional groups. On addition of the copolymer the reaction mixture temperature rose to 68.degree. C. Additional heat was applied to obtain reflux of the reaction mixture and this temperature was maintained for six hours. The

reaction mixture was then allowed to cool to room temperature and was allowed to stand overnight. The solids were then removed by filtration and washed consecutively with methanol, water, methanol, chloroform, benzene, and ether. The resin was then dried in a rotary evaporator at 50.degree. C. The yield of dry solids was 60.8 g. and they were found to have the following chemical analysis, on a weight basis: 3.86% of phosphorus, 13.4% of chlorine, 75.65% of carbon, and 6.19% of hydrogen.

Detailed Description Text (82):

A complex was prepared by dissolving 40 g. of nickel chloride hexahydrate, $\text{NiCl} \cdot 6\text{H}_2\text{O}$, in 1 liter (1.) of acetic acid and 1.5 l. of 1-butanol. The solution was heated to the reflux temperature and to this solution were added 20 g. of the phosphine resin of Example 9. The reaction mixture was stirred and maintained at the reflux temperature for 90 minutes. The complex was isolated by filtration, washed with 1-butanol, washed with methanol, washed with ether, and dried in a rotary evaporator. The yield of complex was 21.3 g. and had a chemical analysis, on a weight basis: 0.44% of Ni, 8.45% of Cl, 3.68% of P, 76.73% of C, and 7.53% of H.

Detailed Description Text (86):

A complex was prepared by dissolving 1.5 g. of bis-benzonitrile-palladium dichloride in 500 ml. of chloroform and adding to this solution about 20 g. of the phosphine resin of Example 9. The reaction mixture was heated to the reflux temperature for 1 hour at which time 50 ml. of benzonitrile were added and reflux maintained for an additional hour. The reaction mixture was then cooled to room temperature. The complex was isolated by filtration, washed with chloroform and dried in a rotary evaporator. The yield of complex was 21.9 g. and had a chemical analysis, on a weight basis: 1.53% of Pd, 11.2% of Cl, 3.9% of P, 70.02% of C, and 4.84% of H.

Detailed Description Text (90):

A complex was prepared by dissolving 1.15 g. of tris-triphenylphosphine-rhodium(I) chloride in 75 ml. of ethanol and adding to this solution 3.0 g. of the phosphine resin of Example 9. The mixture was heated to the reflux temperature for three hours and then cooled to room temperature. The complex was isolated by filtration, washed with ethanol, washed with ether, and dried at 110.degree. C.

Detailed Description Text (95):

The reaction mixture was stirred, heated to the reflux temperature for 5 hours and then cooled to room temperature. The complex was isolated by filtration, washed with water, washed with methanol, washed with ether, and dried. The yield of complex was 10.5 g. and had a chemical analysis, on a weight basis: 2.81% of Rh and 12.2% of S, and 0.93% of Cl.

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L24: Entry 1 of 3

File: USPT

Apr 1, 2003

DOCUMENT-IDENTIFIER: US 6541273 B1

TITLE: Multiple sorbent cartridges for solid phase extraction

Detailed Description Text (12):

Other pesticides which can be evaluated using the methods described herein include 1,2,4-triazinones, 1,3,5-triazines, 2-4-arylenephenoxy propionic acids, 2,6-dinitroanilines, 2-dimethylaminopropane-1,3-dithiol, 2-dimethylaminopropane-1,3-dithiol analogues, Acari, alkanamides, alkylenebis (dithiocarbamates), amides, amidines, aminoglycoside antibiotics, anilides, anilinopyrimidines, antibiotics, aromatic hydrocarbons, aromatic hydrocarbon derivatives, arylalanines, aryloxyalkanoic acids, aryloxyphenoxypropionic acids, auxin, avermectin, azoles,azole analogues, azomethine, bacterium, benzilate, benzimidazoles, benzimidazole precursors, benzofuranyl alkanesulfonates, benzoic acid derivatives, especially chlorinated derivatives, benzonitriles, benzoxazines, benzoylureas, bipyridyliums, bis-carbamates, carbamates, carboxamides, chloracetanilides, chloroamides, chlorophenols, Coleoptera, coumarin anticoagulants, coumarin anticoagulant analogues, cyclodiene organochlorines, cyclohexanedione oximes, cytokinins, diacylhydrazines, dicarboximides, imethylaminopropane-1,3-dithiol, dimethyldithiocarbamate, dinitroaniline, dinitrophenol, dinitrophenol derivatives, diphenyl ethers, Diptera, dithiocarbamates, DMI fungicides, entomopathogenic fungus, ethylene generators, fatty acids, fungi, granulosis viruses, guanidine, halogenated alkanolic acids, Hemiptera, hydroxybenzonitrile, hydroxybenzonitrile precursors, Hymenoptera, imidazolinones, indandione anticoagulants, inorganic compounds, for example, inorganic flourides, juvenile hormone mimics, methyl isothiocyanate precursors, morphactin, morpholine, morpholine analogues (piperidine derivatives), N-trihalomethylthio ompounds, natural pyrethrins, nematodes, Neuroptera, non-ester pyrethroids, nuclear polyhedrosis viruses, organoarsenic compounds, organochlorines, organophosphate esters, organophosphorus compounds, in particular, organophosphorus herbicides, organotin compounds, oxime carbamates, oxyacetamide, phenoxyquinoline, phenyl carbamates, phenylamides (in particular, the acylalanine and acylamino butyrolactone types), phenylamides (in particular, the acylamino coazolidinone type), phenylpyrazoles, phenylpyrroles, phenylureas, pheromones, phosphinico amino acids, phosphoramidates, phthalimides, phthalimide analogues, pyrazoles (acaricides), pyrazole (acaricide) analogues, pyrazoles (herbicides), pyrethroids, pyridazinone (including CBI and PSII), pyridine, pyridinecarboxylic acid, pyrimidine, pyrimidinyl carbinol, pyrimidinylxybenzoic acids, pyrimidinylxybenzoic acid analogues, quaternary ammonium salts, quinolinecarboxylic acids, strobilurin analogues, sulfamoylureas, sulfonylureas, synthetic auxins, tetrazines, thiocarbamates, triazines, triazinones, triazoles, triazolopyrimidine sulfonanilides, triketones, uracil, ureas, especially halogenated ureas, viruses, and wasps.

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L27: Entry 1 of 8

File: USPT

Dec 28, 1999

DOCUMENT-IDENTIFIER: US 6008349 A

TITLE: Melamine polymetaphosphate and process for its production

Brief Summary Text (8):

It is known to employ urea as a condensing agent for the production of a polyphosphate. For example, JP-B-53-2170 discloses a process for producing an amide polyphosphate containing amide-type nitrogen, wherein a phosphoric acid source, such as ammonium orthophosphate, orthophosphoric acid, condensed phosphoric acid, phosphoric anhydride, urea phosphate or a mixture thereof, and a nitrogen source, such as melamine, a cyanamide derivative such as dicyan cyanamide, guanidine or guanyl urea, or a mixture thereof, are subjected to a heat condensation reaction in the presence of a condensing agent, such as urea, urea phosphate or a mixture thereof. With respect to the production conditions, it is disclosed that heat condensation is carried out at a molar ratio of urea/phosphoric acid (as H.sub.3 PO.sub.4)/cyanamide derivative=0.8 to 1.5/1/0.05 to 1 in an atmosphere of ammonia gas at a temperature of from 150 to 350.degree. C. for from 10 minutes to 5 hours, preferably from 1 to 4 hours. Example 1 of the same publication discloses that ammonium phosphate, industrial urea and melamine were mixed at a molar ratio of urea/phosphoric acid (as H.sub.3 PO.sub.4)/melamine=1/1/0.5 to obtain a material for calcination, which was subjected to heat condensation at a temperature of from 240 to 260.degree. C. for 2.5 hours to obtain a substantially water-insoluble amide polyphosphate (nitrogen: 33.48 wt %, diphosphorus pentoxide: 44.73 wt %, elution ratio (10 g of a sample was put into 100 ml of water and stirred for 1 hour at a predetermined temperature, whereupon the elution ratio (%) was determined): 1.36%, pH of the 1% solution: 8.26).

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